and Br III of sodium. Once again the derivative values of the Br II set are low as expected. The Br I derivative data again is low but seems to run more or less parallel to the newer data in a manner that suggests a constant error.

Figure 3a presents the Swenson data at 4·2°K and represents a much better fit of the points to the curve than does the lithium at this temperature.

An analysis of the rather extensive set of sodium data run at various temperatures by Beecroft and Swenson was also run.[17] This study will be presented in the near future.

The potassium data

Potassium presents a particularly difficult problem. In Fig. 4 is presented the Bridgman data. Up to about 60,000 kg/cm² Br II and Br III points seem to present a single set of data. Br I seems to

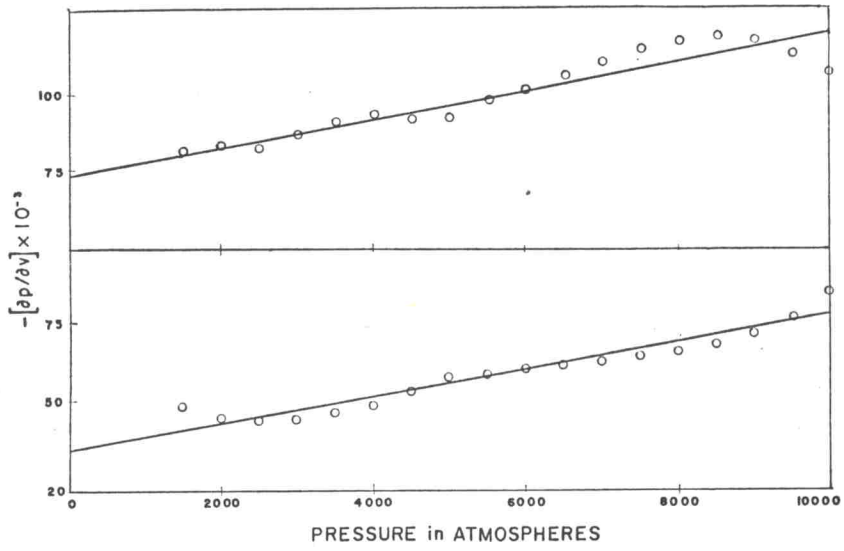

FIG. 3a. Plot of $-(\partial p/\partial v)_T$ vs. pressure in atm. for the Swenson sodium at 4·2°K; ○ Swenson experimental values. —— least squares line obtained from the Swenson experimental values. 3b. Plot of $-(\partial p/\partial v)_T$ vs. pressure in atm. for the Swenson potassium data at 4·2°K. ○ Swenson experimental values.—— least squares line obtained from the Swenson experimental values.

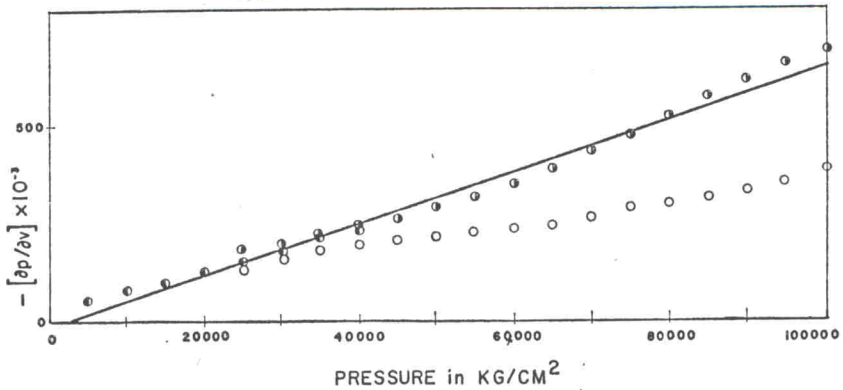

FIG. 4. Plot of $-(\partial p/\partial v)_T$ vs. pressure in kg/cm² for the Bridgman potassium data at room temperature. ◑ Bridgman III experimental values, ◐ Bridgman II experimental values and ○ Bridgman I experimental values. —— least squares line obtained by using the Br II and Br III potassium data.